CO-ORDINATED SCIENCES

Paper 2 Multiple Choice (Extended)

0654/21 May/June 2018 45 minutes

Additional Materials: Multiple Choice Answer Sheet Soft clean eraser Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid. Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you. DO **NOT** WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

Choose the one you consider correct and record your choice in soft pencil on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer. Any rough working should be done in this booklet. A copy of the Periodic Table is printed on page 16. Electronic calculators may be used.

This document consists of 16 printed pages.

1 Which rows correctly match characteristics of living things with their descriptions?

	characteristic	description
1	excretion	removing the waste products of metabolism
2	growth	making more living things of the same type
3	nutrition	taking in or producing food
4	respiration	releasing energy from food

- **A** 1, 2 and 4 **B** 1, 3 and 4 **C** 1 and 3 only **D** 2 and 4 only
- 2 Which statement about cells is correct?
 - A Cell membranes are found only in animal cells.
 - **B** Cell membranes are found only in plant cells.
 - **C** Cell walls are found only in animal cells.
 - D Cell walls are found only in plant cells.
- **3** Which graph shows the effect of light intensity on the rate of photosynthesis, if all other factors are kept constant?



- 4 What leads to coronary heart disease?
 - A Coronary arteries become blocked.
 - **B** Coronary arteries become enlarged.
 - **C** Heart muscles become enlarged.
 - D Heart muscles do not contract.

5 In which tissue does translocation occur and what is a substance that is translocated?

	tissue	substance translocated
Α	phloem	amino acid
В	phloem	glycogen
С	xylem	sucrose
D	xylem	water

6 The word equation for aerobic respiration is shown.

oxygen + \rightarrow carbon dioxide + water

Which molecule is missing from the equation?

- A glucose
- B glycogen
- C starch
- D sucrose
- 7 The diagram shows structures in a section through the front of the eye.



When reading a book, how are the labelled structures involved in focusing the eye?

	1	2	3
Α	contracts	thicker	slackens
в	contracts	thinner	tightens
С	relaxes	thicker	tightens
D	relaxes	thinner	slackens

8 The diagram shows a section through the skin of a person who is sweating.



What happens to the arteriole and what will be the effect on heat loss when a person is sweating?

	arteriole	heat loss
Α	vasoconstricts	decreases
В	vasoconstricts	increases
С	vasodilates	decreases
D	vasodilates	increases

- 9 Which statement about flowers is correct?
 - A The anther and stigma are parts of the carpel.
 - **B** The anther and stigma are parts of the stamen.
 - **C** The ovary and stigma are parts of the carpel.
 - **D** The ovary and stigma are parts of the stamen.

10 The diagram shows a cell that is about to divide by meiosis.



Which cell could be the result of this division?



- 11 What is the function of mitosis?
 - A to produce cells with double the number of chromosomes
 - **B** to produce cells with varying numbers of chromosomes
 - **C** to produce gametes
 - **D** to produce genetically identical cells
- 12 Which processes change the amount of carbon dioxide in the air?

	process causing increase in carbon dioxide	process causing decrease in carbon dioxide
Α	burning fossil fuels	photosynthesis in plants
В	photosynthesis in plants	respiration in animals
С	respiration in animals	respiration in plants
D	respiration in plants	burning fossil fuels

- 13 What is the overuse of nitrogen-containing fertilisers most likely to cause?
 - A acid rain
 - B deforestation
 - **C** eutrophication
 - D global warming
- 14 Which statement about liquids is correct?
 - A They have a fixed shape and a fixed volume.
 - **B** They have a fixed shape but not a fixed volume.
 - **C** They have no fixed shape but they do have a fixed volume.
 - **D** They have no fixed shape and no fixed volume.
- **15** Pure copper chloride can be obtained from a mixture of powdered copper and solid copper chloride.

Three stages in the method are listed.

- P add water and stir
- Q crystallise
- R filter

In which order are these stages carried out in order to obtain pure copper chloride from the mixture?

- $\textbf{A} \quad \textbf{P} \ \rightarrow \ \textbf{Q} \ \rightarrow \ \textbf{R}$
- $\textbf{B} \quad \textbf{P} \rightarrow \textbf{R} \rightarrow \textbf{Q}$
- $\textbf{C} \quad \textbf{R} \rightarrow \textbf{P} \rightarrow \textbf{Q}$
- $\textbf{D} \quad \textbf{R} \, \rightarrow \, \textbf{Q} \, \rightarrow \, \textbf{P}$

16 The atomic number of potassium is 19.

What is the electronic structure of a potassium atom?



17 A rock contains three ores, galena (PbS), copper pyrites (CuFeS₂) and cinnabar (HgS).How many metals are present in this rock?

A 3 **B** 4 **C** 5 **D** 8

18 What is the equation for the complete combustion of ethane?

$$\mathbf{A} \quad \mathsf{C}_2\mathsf{H}_6 \ + \ 2\mathsf{O}_2 \ \rightarrow \ 2\mathsf{CO}_2 \ + \ 3\mathsf{H}_2\mathsf{O}$$

- $\textbf{B} \quad 2C_2H_6 \ \textbf{+} \ 3O_2 \ \rightarrow \ 4C \ \textbf{+} \ \ 6H_2O$
- $\label{eq:constraint} \begin{array}{ccc} \mbox{C} & 2C_2H_6 \mbox{ + } 5O_2 \mbox{ \rightarrow } 4CO \mbox{ + } 6H_2O \end{array}$
- $\textbf{D} \quad 2C_2H_6 \ \textbf{+} \ 7O_2 \ \rightarrow \ 4CO_2 \ \textbf{+} \ \ 6H_2O$

- **19** What are the products of the electrolysis of concentrated aqueous sodium chloride?
 - A chlorine, hydrogen and sodium
 - B chlorine, hydrogen and sodium hydroxide
 - **C** chlorine and hydrogen only
 - D chlorine and sodium only
- 20 Which statement describes an exothermic reaction?
 - **A** Heat energy is transferred from the surroundings and the temperature decreases.
 - **B** Heat energy is transferred from the surroundings and the temperature increases.
 - **C** Heat energy is transferred to the surroundings and the temperature decreases.
 - **D** Heat energy is transferred to the surroundings and the temperature increases.
- **21** Dilute hydrochloric acid is added to lumps of calcium carbonate.

Which change decreases the rate of the reaction?

- A Decrease the temperature of the acid.
- **B** Increase the concentration of the acid.
- **C** Use a larger volume of the acid.
- **D** Use powdered calcium carbonate.
- 22 When iron is heated with steam, a black solid is formed.



The equation for the reaction is shown.

iron + water \rightarrow iron oxide + hydrogen

Which statement about this reaction is correct?

- A Iron has been oxidised because it has gained oxygen.
- **B** Iron has been reduced because it removed oxygen from water.
- **C** Iron oxide has been reduced because it contains oxygen.
- **D** Water has been oxidised because it contains oxygen.

23 Zinc sulfate is made by adding zinc oxide to dilute sulfuric acid.

The steps used to obtain zinc sulfate crystals are listed.

- 1 filter the solution to remove excess zinc oxide
- 2 warm the zinc sulfate solution
- 3 add excess zinc oxide and stir
- 4 filter and dry the crystals

What is the correct order of the steps?

- $\textbf{A} \quad 1 \rightarrow 3 \rightarrow 4 \rightarrow 2$
- $\textbf{B} \quad 2 \rightarrow 1 \rightarrow 3 \rightarrow 4$
- $\textbf{C} \quad 3 \rightarrow 1 \rightarrow 2 \rightarrow 4$
- $\textbf{D} \quad 3 \rightarrow 2 \rightarrow 1 \rightarrow 4$
- 24 Which diagram represents the structure of an alloy?



25 Four iron nails are placed in four test-tubes as shown.

In which test-tube does the iron nail rust most quickly?



26 During the manufacture of sulfuric acid by the Contact process, sulfur trioxide is produced.

The sulfur trioxide is dissolved in concentrated sulfuric acid.

Which statement explains why sulfur trioxide is not dissolved in water?

- **A** The reaction is too endothermic.
- **B** The reaction is too exothermic.
- **C** The reaction is too slow.
- **D** The reaction needs a high pressure.
- 27 Ethanol is manufactured by reacting ethene with steam in the presence of a catalyst.

Which type of reaction occurs?

- A addition
- **B** oxidation
- **C** polymerisation
- D reduction
- **28** A body has a mass of 12 kg and weighs 120 N on Earth. It is taken from Earth to the Moon.

The strength of the gravitational field on the Moon is one sixth of that on Earth.

What is the mass and what is the weight of the body on the Moon?

	mass/kg	weight/N
Α	2.0	20
в	2.0	120
С	12	20
D	12	120

29 The diagrams show four solid blocks with the same mass.

Which block is made from the least dense material?



30 A rocket has a mass of 300 kg. Its motors produce a force of 12000 N vertically upwards. The acceleration of free fall g is 10 m/s^2 .

What is the resultant force on the rocket and what is the acceleration of the rocket?

	resultant force/N	acceleration m/s ²
Α	9000	30
В	9000	2.7 × 10 ⁶
С	15000	50
D	15000	4.5×10^{6}

31 The speed-time graph represents the journey of a bicycle.



32 A glass bottle containing warm air is sealed with a screw cap and then cooled in cold water.



The contraction of the glass bottle can be ignored.

What remains the same during the cooling?

- the air pressure inside the bottle Α
- В the energy of the air molecules in the bottle
- С the force on the cap made by the air molecules in the bottle
- D the volume of air in the bottle
- 33 When a substance changes state, it releases latent heat of fusion.

What is the change of state?

- Α gas to liquid
- В liquid to gas
- С liquid to solid
- solid to liquid D

Α

34 A wave passes from medium 1 into medium 2. The diagram shows the change in direction of the wave.



How do the frequency and the wavelength of the wave change, if at all, as it passes from medium 1 into medium 2?

	frequency	wavelength
Α	decreases	decreases
В	decreases	increases
С	no change	decreases
D	no change	increases

35 Light travelling in a glass block strikes the inside surface of the block at the critical angle.

What is the size of the angle of refraction?

- **A** equal to the critical angle
- **B** between the critical angle and 90°
- **C** exactly 90°
- D greater than 90°

36 A bar magnet is brought near to a metal rod.



The magnet is now turned around so that the N-pole is on the right. The magnet is again brought near to the metal rod.

In both cases the metal rod is attracted to the magnet.

What could the metal rod be?

- A another bar magnet
- **B** a piece of aluminium
- **C** a piece of copper
- D a piece of iron
- **37** Which quantity is defined in terms of the energy supplied by a source in driving charge round a complete circuit, and what is its unit?

	quantity	unit
Α	e.m.f.	joule
В	e.m.f.	volt
С	p.d.	joule
D	p.d.	volt

38 The diagrams show four circuits.

Which circuit contains two lamps connected in parallel with each other, and contains an ammeter that measures the total current in the two lamps?



39 Transmission cables are used to carry electricity between a power station and a town.

Near the power station a transformer is used to reduce energy losses in the transmission cables.



How do the voltage of the transmission cables and the current in them compare with their values for the power station cables?

	transmission cable voltage	transmission cable current
Α	larger	larger
В	larger	smaller
С	smaller	larger
D	smaller	smaller

40 The diagram shows a beam of α -particles and γ -rays entering an electric field between two metal plates.



What is the effect, if any, of the electric field on the α -particles and on the γ -rays?

	α -particles	γ-rays
Α	deflected downwards	deflected downwards
В	deflected downwards	not deflected
С	deflected upwards	deflected downwards
D	deflected upwards	not deflected

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The Periodic Table of Elements

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	>	7	He	heliun 4	10	Ne	neon 20	18	Ar	argor 40	36	Ϋ́	krypto o 4	5 1	54	Xe	xenor 131	86	Rn	radon.	1				
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2	2				9	U	carbon 12	14	Si	silicon 28	32	Ge	germanium	2	20	Sn	tin 119	82	Pb	lead	207	114	Fl	flerovium	I
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											29	Cu	copper	5 !	47	Ag	silver 108	79	Au	gold	197	111	Rg	roentgenium	I
dno										28	ïZ	nickel	60	46	Pd	palladium 106	78	Ţ	platinum	195	110	Ds	darmstadtium	I	
GG					_						27	ပိ	cobalt E0		45	Rh	rhodium 103	77	L	iridium	192	109	Mt	meitnerium	I
			т	hydrogen 1							26	Fе	iron E.e.	8	44	Ru	ruthenium 101	76	SO	osmium	190	108	Hs	hassium	I
										25	Mn	manganese	3	43	ц	technetium -	75	Re	rhenium	186	107	Bh	bohrium	I	
						bol	SS				24	ŗ	chromium	3	42	Мо	molybdenum 96	74	\geq	tungsten	184	106	Sg	seaborgium	I
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											21	လိ	scandium	<u></u>	39	≻	yttrium 89	57-71	lanthanoids			89-103	actinoids		
=	=				4	Be	beryllium 9	12	Mg	magnesium 24	20	Ca	calcium	P 6	38	ي ک	strontium 88	56	Ba	barium	137	88	Ra	radium	I
_	_				ю	:	lithium 7	11	Na	sodium 23	19	¥	potassium	0	37	Rb	rubidium 85	55	Cs	caesium	133	87	Ļ	francium	I

	/9	58	66	60	61	29	63	64	69	66	67	68	69	07	17
lanthanoids	La	Ce	Pr	Νd	Рm	Sm	Еu	Вd	Tb	D	Ю	ц	Tm	Υb	Lu
	lanthanum 139	cerium 140	praseodymium 141	neodymium 144	promethium –	samarium 150	europium 152	gadolinium 157	terbium 159	dysprosium 163	holmium 165	erbium 167	thulium 169	ytterbium 173	lutetium 175
	89	06	91	92	93	94	95	96	97	98	66	100	101	102	103
actinoids	Ac	Тh	Ра		dN	Pu	Am	Cm	剐	Ç	Es	Еm	Md	No	L
	actinium	thorium	protactinium	uranium	neptunium	plutonium	americium	curium	berkelium	californium	einsteinium	fermium	mendelevium	nobelium	lawrencium
	I	232	231	238	I	I	I	I	I	I	I	I	I	ļ	I

The volume of one mole of any gas is $24\,dm^3$ at room temperature and pressure (r.t.p.).

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